CHEMICAL REACTIONS – SET IV

You are given 10 aqueous solutions [Li_3PO_4, $AgNO_3$, KNO_3, $Cu(NO_3)_2$, NaI, Na_2CO_3, K_2SO_4, NaOH, HCl, and K_2CrO_4] , two solids (Mg and Cu), and unlimited distilled water. Write a balanced equation for an actual reaction that would produce the following. Note: There may be more than one product formed.

(a) solid silver iodide

(b) liquid water

(c) hydrogen gas

(d) carbon dioxide gas

(e) solid silver chromate

(f) aqueous magnesium chloride

(g) solid copper (II) hydroxide